

Answers.

Dissociation Equation	K_{sp} expression
$\text{AlPO}_4 \rightleftharpoons \text{Al}^{3+}(\text{aq}) + \text{PO}_4^{3-}(\text{aq})$	$K_{sp} = [\text{Al}^{3+}][\text{PO}_4^{3-}]$
$\text{BaSO}_4 \rightleftharpoons \text{Ba}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$	$K_{sp} = [\text{Ba}^{2+}][\text{SO}_4^{2-}]$
$\text{CdS} \rightleftharpoons \text{Cd}^{2+}(\text{aq}) + \text{S}^{2-}(\text{aq})$	$K_{sp} = [\text{Cd}^{2+}][\text{S}^{2-}]$
$\text{Cu}_3(\text{PO}_4)_2 \rightleftharpoons 3 \text{Cu}^{2+}(\text{aq}) + 2 \text{PO}_4^{3-}(\text{aq})$	$K_{sp} = [\text{Cu}^{2+}]^3 [\text{PO}_4^{3-}]^2$
$\text{CuSCN} \rightleftharpoons \text{Cu}^+(\text{aq}) + \text{SCN}^-(\text{aq})$	$K_{sp} = [\text{Cu}^+][\text{SCN}^-]$
$\text{Hg}_2\text{Br}_2 \rightleftharpoons \text{Hg}_2^{2+}(\text{aq}) + 2 \text{Br}^-(\text{aq})$	$K_{sp} = [\text{Hg}_2^{2+}][\text{Br}^-]^2$
$\text{AgCN} \rightleftharpoons \text{Ag}^+(\text{aq}) + \text{CN}^-(\text{aq})$	$K_{sp} = [\text{Ag}^+][\text{CN}^-]$
$\text{Zn}_3(\text{AsO}_4)_2 \rightleftharpoons 3 \text{Zn}^{2+}(\text{aq}) + 2 \text{AsO}_4^{3-}(\text{aq})$	$K_{sp} = [\text{Zn}^{2+}]^3 [\text{AsO}_4^{3-}]^2$
$\text{Mn}(\text{IO}_3)_2 \rightleftharpoons \text{Mn}^{2+}(\text{aq}) + 2 \text{IO}_3^-(\text{aq})$	$K_{sp} = [\text{Mn}^{2+}][\text{IO}_3^-]^2$
$\text{PbBr}_2 \rightleftharpoons \text{Pb}^{2+}(\text{aq}) + 2 \text{Br}^-(\text{aq})$	$K_{sp} = [\text{Pb}^{2+}][\text{Br}^-]^2$
$\text{SrCO}_3 \rightleftharpoons \text{Sr}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq})$	$K_{sp} = [\text{Sr}^{2+}][\text{CO}_3^{2-}]$
$\text{Bi}_2\text{S}_3 \rightleftharpoons 2 \text{Bi}^{3+}(\text{aq}) + 3 \text{S}^{2-}(\text{aq})$	$K_{sp} = [\text{Bi}^{3+}]^2 [\text{S}^{2-}]^3$